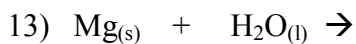
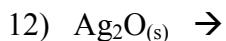
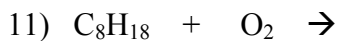
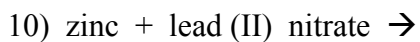
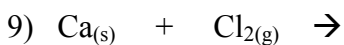
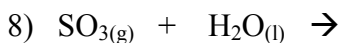
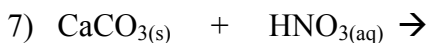
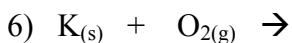
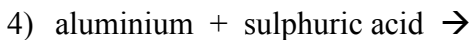
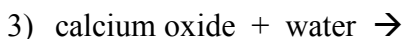
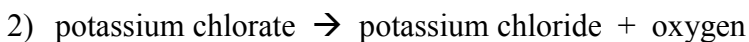


## Chemical Reactions Worksheet

For each of the following equations, determine the products if necessary, write a balanced equation, classify what kind of chemical reaction it is (addition, decomposition, single displacement...etc.) and for single or double displacement reactions, also include the net ionic equation.



15) perchloric acid + calcium hydroxide →

16) sodium hydroxide + phosphoric acid →

17)  $\text{K}_2\text{CO}_{3(\text{aq})}$  +  $\text{HNO}_{3(\text{aq})}$  →

18)  $\text{FeBr}_{3(\text{s})}$  →

19)  $\text{FeCl}_{3(\text{aq})}$  +  $\text{Na}_2(\text{SO}_4)_{(\text{aq})}$  →

20)  $\text{Ba}(\text{NO}_3)_{2(\text{aq})}$  +  $\text{Mg}(\text{SO}_4)_{(\text{aq})}$  →

21)  $\text{Cl}_{2(\text{g})}$  +  $\text{MgBr}_{2(\text{aq})}$  →

22) barium + water →

23)  $\text{CH}_4$  +  $\text{O}_2$  →

24) zinc + sodium phosphate →

25) copper (II) chloride + potassium phosphate →

26) calcium carbonate (solid) →

27) lithium oxide + carbon dioxide →

28) ammonium sulphate + magnesium chloride →

29) calcium nitrate + copper (I) carbonate →

30) copper (II) sulphate pentahydrate →